CHROM. 4193

Poly-N-vinyl pyrrolidone chromatography

Effect of pH on elution of bases and a proposed mechanism for the adsorption of purines

Recently the separation of purine and pyrimidine bases by poly-N-vinyl pyrrolidone (PVP) chromatography was reported[1]. The principal mechanism of adsorption was attributed to hydrogen and hydrophobic bonding. Electrolytes tend to strengthen hydrophobic bonds whereas they weaken salt linkages[2]. Adenine in distilled water elutes at approximately 80 ml, while adenine in 0.3 ml saturated $(NH_4)_2SO_4$ elutes at 56 ml (1 × 39.5 cm column). The elution volume of guanine is also decreased in salt[3]. These observations seem to indicate that hydrophobic bonds are not of major importance in the adsorption of adenine or guanine to PVP.

Compared with the pyrimidines, adenine and guanine bind tenaciously to PVP. Our present studies, reported below, indicate that hydrogen bonding increases the electrophilic character of the carbonyl carbon of PVP and hence the center of high electron density in adenine and guanine adjacent to the point of hydrogen bonding may also be involved.

Materials and methods

Adenine, guanine, cytosine and thymine were purchased from Nutritional Biochemical Company* and 1-methylguanine from Sigma Chemical Company. Insoluble PVP (sold under the trade name of Polyclar AT Powder) was obtained from GAF Corporation, New York.

The Polyclar AT was mixed with distilled water, and the fines were discarded by repeated decantation. The suspension was poured into a column and allowed to pack with gravity flow.

Columns were equilibrated with the eluting solution before standard samples were added. A 0.5 ml solution containing 0.08 mg each of thymine, cytosine, guanine and adenine (adjusted to the pH of the eluting medium) was applied to a 1.0 × 39.5 cm column. Elution was carried out at room temperature and atmospheric pressure with solutions containing 0.02 *M* NH_4Cl + 0.2 *M* NH_4OH at pH 10.3 and pH 7 or 0.005 *N* HCl adjusted to pH 3.5 with dilute NH_4OH. 1-Methylguanine (0.1 mg in 0.5 ml) was eluted at pH 7.

The effluent was collected in 1.25 ml fractions and monitored at 260 mμ in a Gilford 220 spectrophotometer. Identification of the bases was made from the ultraviolet absorbance spectra.

Results and discussion

The dissociations of adenine and guanine of p*K'* values 4.2 and 3.3, respectively, may be attributed to the amino groups and the guanine dissociation of p*K'* 9.6 to the $-N_1H-C_6O-$ group. The dissociation of adenine of p*K'* 9.8 and the dissociation of guanine of p*K'* 12.3 are almost certainly the dissociation of –NH– in the imidazole ring[4].

* Mention of trade or company names does not imply endorsement by the Department over others not named.

TABLE I

ELUTION OF BASES WITH VARIATION IN pH

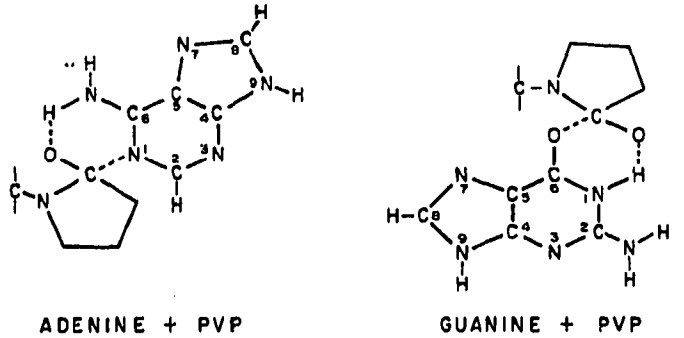
Dissociation constants (pK') are taken from JORDAN[4].

	pK'			Ve (ml)[a]		
	$-NH_3^+$	$-N_1H-C_6O-$	*Imidazole* $-NH-$	*pH 3.5*	*pH 7*	*pH 10.3*
Adenine	4.2		9.8	43.8	73.8	70.0
Guanine	3.3	9.6	12.3	50.0	67.5	37.5
Cytosine	4.6			31.2	34	32.5
Thymine		9.8		36.2		28.8
1-Methylguanine					52.5	

[a] Ve = elution volume.

Table I shows the variation in elution volume (Ve) of adenine, guanine, cytosine and thymine under different elution conditions. There appears to be little contribution from the imidazole –NH– of adenine, since there is only a slight difference between the elution volumes at pH 7 and pH 10.3 (74 *vs.* 70 ml). The only large change in elution volume of adenine occurred at pH 3.5 (44 ml). Hence the amino group of adenine (pK' 4.2) must play a significant role in binding to PVP. Hydrogen bonding alone could not account for the greater attraction of adenine to PVP compared with that of the pyrimidines, which have the same number of possible hydrogen bonding sites. Because of the electrophilic nature of the pyrrolidone carbonyl group, it is reasonable to implicate the high electron density at $N_{(1)}$ of adenine as a site of adsorption. With hydrogen bonding of the amino group, the electrophilic character of the carbonyl carbon is increased[5], and the center of high electron density at $N_{(1)}$ of adenine is attracted to the carbonyl carbon of PVP, making ring formation possible as indicated in Fig. 1. When the hydrogen bond involving the adenine amino group is ruptured at pH 3.5 the ring is broken and adenine elutes much sooner.

Similarly, guanine could bind to PVP through formation of a six-membered ring. Once the $-N_1H-$ of guanine is hydrogen bonded, the guanine keto oxygen may be attracted to the electrophilic center of the PVP molecule (Fig. 1). At pH 10.3 guanine has a much smaller elution volume (37.5 ml) than at pH 7 (67.5 ml). This same 30 ml change (as indicated above with adenine) can be attributed to the



ADENINE + PVP GUANINE + PVP

Fig. 1. Postulated adsorption of adenine and guanine to PVP.

ionization of $-N_1H-$ (pK' 9.6) disrupting the hydrogen bonding and hence the ring.

The *Ve* of 1-methylguanine (pH 7) is less than that of guanine. Substitution of the hydrogen atom at $N_{(1)}$ of guanine with a methyl group leads to a decrease in adsorption, presumably due to lack of ring formation.

Hence, although hydrogen bonding has been considered the principal mechanism by which substances adsorb to PVP, the hydrogen bonding may be concerted with a nucleophilic attack at the pyrrolidone carbonyl carbon, resulting in ring formation and increased adsorption. The feasibility of the stereo-structural relationships shown in Fig. 1 has been confirmed by use of molecular models.

As a practical application of altering the elution pH, cytosine and thymine can be separated by PVP chromatography at pH 3.5, since the $-N_1H-$ of thymine is hydrogen bonded to PVP at this pH and elutes at 36 ml. The two bases may also be resolved at pH 10.3 where the $-N_1H-$ of thymine is ionized (pK' 9.8), disrupting the hydrogen bonding and allowing thymine to be eluted first.

The authors wish to thank Mr. L. BLECHER of GAF Corporation for providing the Polyclar AT Powder.

Eastern Utilization Research and Development Division,
Agricultural Research Service,
United States Department of Agriculture,
Philadelphia, Pa. 19118 (U.S.A.)

THERESE M. DOUGHERTY
ABNER I. SCHEPARTZ

1 J. LERNER, T. M. DOUGHERTY AND A. I. SCHEPARTZ, *J. Chromatog.*, 37 (1968) 453.
2 W. KAUZMANN, *Advan. Protein Chem.*, 14 (1959) 49.
3 T. M. DOUGHERTY AND A. I. SCHEPARTZ, *J. Chromatog.*, 40 (1969) 299.
4 D. O. JORDAN, *The Chemistry of Nucleic Acids*, Butterworths, Washington, 1960, p. 134.
5 E. H. CORDES AND W. P. JENCKS, *J. Am. Chem. Soc.*, 84 (1962) 4319.

Received May 20th, 1969

CHROM. 4192

Behavior of uronic acids and acid-treated uronic acids on an automatic amino acid analyzer

SCHRAM *et al.*[1] reported briefly that ninhydrin reacted with non-nitrogen-containing carbohydrate material in a protein hydrolysate. More recently, the reaction of ninhydrin with non-nitrogenous compounds such as carbohydrates and related compounds has been investigated by ZACHARIUS AND TALLEY[2], ZACHARIUS AND PORTER[3], and SCHILLING *et al.*[4,5].

The kinds of compounds reported to give positive peaks on an amino acid analyzer include hydroxy acids, keto acids, hydroxy aldehydes, keto aldehydes, aldonic and uronic acid derivatives, hydroxy ketones, and some monosaccharides. In general, the reaction products from these compounds absorb more at 440 mμ than at 570 mμ. This high 440:570 mμ absorption ratio is similar to that found for the